



CHEMISTRY NMDCAT

(UNIT-9)

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03418729745(WhatsApp Groups)

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TOPICS:-

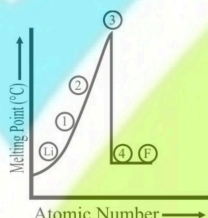
✓ s AND p BLOCK ELEMENTS

- Q.1** Indicate the parameter, which effect atomic Radii in a period
- Number of shells
 - Nuclear charge
 - Shielding effect
 - No. of orbitals
- Q.2** Size of inert gas is _____ than proceeding halogens
- Larger
 - Smaller
 - Equal
 - Variable
- Q.3** Periodic variation of atomic size in periodic table is similar to that off
- Ionic radii
 - Ionization energy
 - Electron Affinity
 - Electronegativity
- Q.4** Select correct increasing order of atomic radius
- Ne > O > S > Al
 - Ne < O > S > Al
 - Ne < O < S < Al
 - Ne > O < S > Al
- Q.5** Cation is smaller in size than parent atom because of
- Greater effective Nuclear charge and lesser shielding effect
 - Greater effective Nuclear charge and greater shielding effect
 - Lesser effective Nuclear charge and lesser shielding effect
 - Greater effective Nuclear charge and greater shielding effect
- Q.6** Indicate incorrect order of atomic and ionic radii
- Na > Na⁺
 - O⁻² < O⁻¹
 - Cl⁻ > Cl
 - Ar > Cl
- Q.7** When an atom is ionized, the electron is removed from _____.
- Inner most shell
 - Outermost shell
 - Second last shell
 - Any shell
- Q.8** Ionization energy of a neutral atom does not depends upon
- Atomic size
 - Nuclear charge
 - Shielding effect
 - Cationic size
- Q.9** Mark the correct statement, ionization energy of an element in a period increases due to
- Successive addition of e⁻ shell
 - Successive increase of nuclear charge
 - Successive increase of effective nuclear charge
 - Both b and c
- Q.10** With respect to 1st ionization energy, abnormal ionization energy trend are shown by _____ and _____ elements of period 3
- Mg and Al
 - Mg and S
 - Al and S
 - P and S

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- Q.25 Mark the correct statement**
 a. Non-metallic character decrease as atomic size increases
 b. Non-metals yield acidic oxides
 c. Non-metals are poor conductor of heat and electricity
 d. All of the above
- Q.26 Melting and boiling points of period 2 vary in order**
 a. Increase
 b. Decrease
 c. 1st Increase and then decrease
 d. 1st decrease and then increase
- Q.27 In group VII-A melting points**
 a. Increase down the group
 b. Decrease down the group
 c. 1st Increase and then decrease
 d. 1st decrease and then increase
- Q.28 The diagram below is a plot of melting points of elements of second period against their atomic numbers. Lithium and fluorine are placed at the extreme ends of the plot. On the basis of melting points where would you place carbon among the empty slots on the plot?**



- a. 1
 b. 4
 c. 2
 d. 3
- Q.29 Electrical conductivity is due to free movement of**
 a. Ions
 b. Electrons
 c. Atoms
 d. Nucleus
- Q.30 Electrical conductivity of elements of period 3 increases from Na to _____**
 a. Mg
 b. Al
 c. P
 d. Ar
- Q.31 Lowest melting point is of**
 a. Be
 b. Mg
 c. Ca
 d. Sr
- Q.32 Alkali metals react with H₂O, they produce metal hydroxides and**
 a. CO₂
 b. H₂
 c. O₂
 d. N₂
- Q.33 Li when burn in O₂ mainly gives**
 a. LiO
 b. LiO₂
 c. Li₂O
 d. Li₂O₂
- Q.34 Which of the following alkali metal form peroxide on reacting with O₂**
 a. Na
 b. K
 c. Rb
 d. Cs
- Q.35 $2\text{NaNO}_3 + 10\text{Na} \rightarrow 6\text{Y} + \text{N}_2$, What is "Y" in the given reaction**
 a. Na₂O₂
 b. NaO
 c. Na₂O
 d. NaO₂
- Q.36 Nitrate of which of the following alkali metal decompose to give oxide**
 a. Li
 b. Na
 c. K
 d. Rb
- Q.37 Guess the product formed in given reaction**



- a. Reactivity of IA with H_2O increase down the group
- b. Reactivity of IIA with H_2O increase down the group
- c. Mg react with H_2O vigorously
- d. Calcium reacts with H_2O to form calcium hydroxide

d. CaO

d. Mg

d. SrO

d. HClO_3

d. Al_2O_4

d. All of above

d. C

d. All of above

d. F

d. Cs

d. Golden yellow

The graph plots the melting point in degrees Celsius against the atomic number for the elements Na, Mg, Al, Si, P, S, Cl, and Ar. The melting point starts at approximately 98°C for Na, rises to 1200°C for Mg, 933°C for Al, and peaks at 1414°C for Si. It then drops sharply to 44°C for P, and continues to decrease to 115°C for S, -101°C for Cl, and -189°C for Ar.

Element	Atomic Number	Melting Point (°C)
Na	11	98
Mg	12	1200
Al	13	933
Si	14	1414
P	15	44
S	16	115
Cl	17	-101
Ar	18	-189

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- a. Decrease in atomic radius from 'Si' to 'P'
- b. Change in bonding and structure of two elements
- c. Different densities of two elements
- d. Increase in electron density from 'Si' to 'P'

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Chem T-9

	A	B	C	D		A	B	C	D		A	B	C	D		A	B	C	D
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